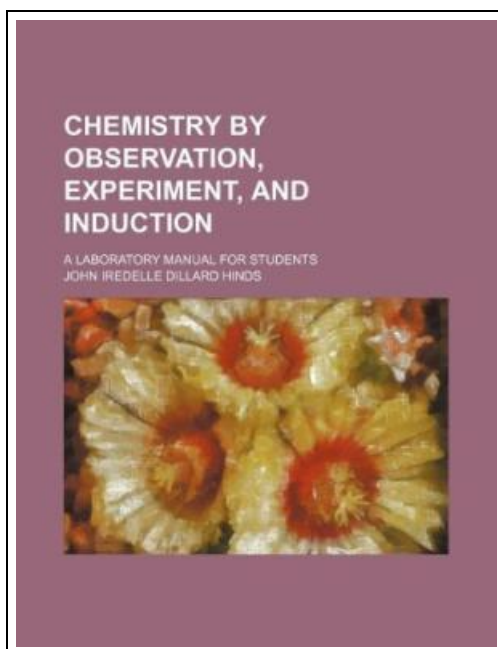


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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1905 Excerpt: .dissociating into NO₂. Write the equation: Test with burning match. Invert the bottle over water. (b) Place in an evolution-flask a few grams of granulated tin and connect the delivery-tube with a U tube immersed in a freezing mixture. Add strong nitric acid through the funnel-tube. Liquid N₂O₄ collects in the U tube. Pour out in a beaker and see how rapidly it evaporates. Properties.--Physical state? Color of liquid? Color of gas? Odor? Combustible? Support combustion? Soluble? Valence of the N? Structural formula? What is formed when it dissolves in water? Write the equation: NITRIC ACID, HNO₃. Preparation.--Place in a small retort 4 g. potassium or sodium nitrate and add 6 cc. strong sulfuric acid. Heat gently and condense the nitric acid in a test-tube or receiver cooled with water. Principle: Same as under hydrogen chlorid. Explain the action: Write the equation in ions: Experiments.--Place a few drops of the acid upon a bit of zinc in a beaker. Repeat the experiment with copper, iron, copper oxid, and calcium carbonate. What are the brown fumes? Place a drop on a piece of colored cloth; on paper. Dilute some of the acid and taste it. Drop a very small bit of phosphorus (large as a pin s head) on the acid. Oxidizing Action.--(a) To 10 cc. of ferrous sulfate solution add a few drops of nitric acid and boil. Ferrous ion colorless is oxidized to ferric ion red. (b) To 1 g. of sugar in a...



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