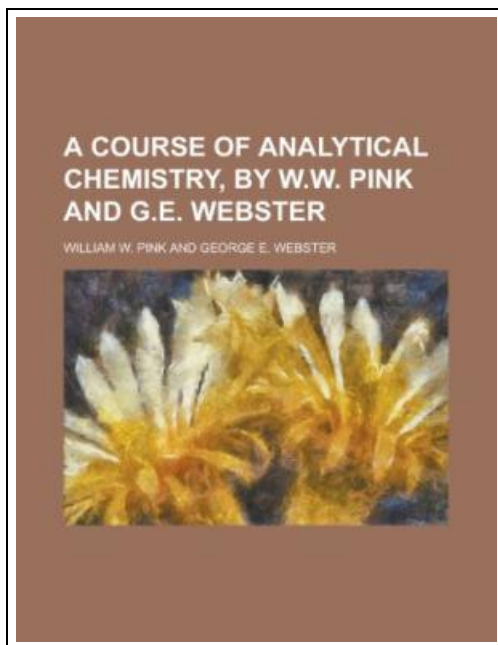


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This book is great. I have go through and so i am confident that i will going to read through once again again in the future. I am just easily can get a satisfaction of looking at a written book.

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Rarebooksclub.com, United States, 2013. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can usually download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1874 edition. Excerpt: . Acid. 121. Add MnO₂, manganic oxide (black oxide of manganese), and SO₂, H₂O, then apply heat; evolution of chlorine takes place, known by its smell, colour, c. 122. NO₂Ag₂O gives a curdy-looking white ppt. of AgCl, argentic chloride (chloride of silver), soluble in AmHo, but insoluble in weak NO₂Ho; exposure to the sunlight always darkens this ppt. HBr.--Hydrobromic Acid. 123. Add MnO₂ and SO₂H₂O, then apply heat; evolution of bromine vapour takes place, which colours moistened starch paper ytlhw. 124. NO₂Ag₂O gives a cream-coloured yellow ppt. of AgBr, slightly soluble in AmHo, but insoluble in NO₂Ho. HI.--Hydriodic Acid. 125. Treat as in Reaction No. 123; evolution of iodine vapour of a violet or purple colour takes place, which colour moistened starch paper blue. 126. NO₂Ag₂O gives a light yellow ppt. of AgI, argentic iodide (iodide of silver), soluble in KCy, but insoluble in weak NO₂Ho. No 127. ij. Hgo, mercuric nitrate (nitrate of mercury), gives a greenish-yellow ppt., insoluble in excess of the reagent. SH₂--Hydrosulphuric Acid. 128. Addition of HCl and application of heat causes evolution of SH₂, known by its nasty smell; sometimes a piece of zinc is required in the case of those sulphides insoluble in HCl. C(Me)O 129. Pbo gives a black ppt. of PlS, plumbic C(Me)O sulphide (sulphide of lead), with soluble sulphides. coio AcEXIC AcID (This acid is only partly precipitated by NO₂Ag₂O.) 129a. Add Fe₂Cl₆, ferric chloride (perchloride of iron), to acetic solutions; it will change them...



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